<table>
<thead>
<tr>
<th><strong>Title</strong></th>
<th>Use of amendments in a peat soil to reduce phosphorus losses from forestry operations</th>
</tr>
</thead>
<tbody>
<tr>
<td><strong>Author(s)</strong></td>
<td>Callery, Oisin; Brennan, Raymond B.; Healy, Mark G.</td>
</tr>
<tr>
<td><strong>Publication Date</strong></td>
<td>2015-10-18</td>
</tr>
<tr>
<td><strong>Publisher</strong></td>
<td>Elsevier</td>
</tr>
<tr>
<td><strong>Link to publisher's version</strong></td>
<td><a href="https://doi.org/10.1016/j.ecoleng.2015.10.016">https://doi.org/10.1016/j.ecoleng.2015.10.016</a></td>
</tr>
<tr>
<td><strong>Item record</strong></td>
<td><a href="http://hdl.handle.net/10379/6671">http://hdl.handle.net/10379/6671</a></td>
</tr>
<tr>
<td><strong>DOI</strong></td>
<td><a href="http://dx.doi.org/10.1016/j.ecoleng.2015.10.016">http://dx.doi.org/10.1016/j.ecoleng.2015.10.016</a></td>
</tr>
</tbody>
</table>

Some rights reserved. For more information, please see the item record link above.
Use of amendments in a peat soil to reduce phosphorus losses from forestry operations

O. Callery, R.B. Brennan*, M.G. Healy

Civil Engineering, National University of Ireland, Galway, Co. Galway, Rep. of Ireland.

*Corresponding author. Tel: +353 91 494442; Fax: +353 91 494507. E-mail address:

raymond.brennan@nuigalway.ie

Abbreviations: Al-WTR, aluminum water treatment residual
Abstract

Forestry harvesting on peats is known to result in significant losses of soil phosphorus (P) to adjacent waters, and the issue is becoming an increasingly serious concern as peatland forest stocks mature and reach harvestable age. One potential solution could be the use of low-cost P recovery techniques based on the chemical precipitation and/or adsorption of the dissolved fraction of soil P, which would otherwise be lost. Such recovery techniques have shown promise in similar applications on mineral soils. However, the interaction of peat with P adsorbing materials can significantly alter their adsorptive characteristics, and it is consequentially not known what materials might be suitable for this application. This study compared the performance of six potential soil amendments (aluminum water treatment residual (Al-WTR), crushed concrete, gypsum, magnesium hydroxide, magnesium oxide, and steel wool) in removing P from aqueous solution in the presence of a typical forest peat soil. Comparison of adsorption isotherms plotted from these batch adsorption studies showed that the observed P adsorption maxima of Al-WTR and steel wool were increased by the presence of peat, from 10.6 mg g\(^{-1}\) and 20.4 mg g\(^{-1}\), to 11.8 mg g\(^{-1}\) and 52.5 mg g\(^{-1}\), respectively. In contrast, the observed P adsorption maxima of crushed concrete, gypsum, and magnesium oxide were reduced in the presence of peat, by 44%, 87%, and 37%, respectively. The maximum P adsorption achieved by magnesium hydroxide was increased from 29.8 mg g\(^{-1}\) to 59 mg g\(^{-1}\) at an amendment to peat-solid ratio of 1:4, but decreased from 73.9 mg g\(^{-1}\) to 23.6 mg g\(^{-1}\) at an amendment to peat-solid ratio of 1:10. It was concluded that Al-WTR, in particular, shows considerable promise for use as a soil amendment for P immobilization in a peat environment.

Keywords: Phosphorus; forestry; peat; clearfelling; harvesting; buffer zones; adsorption
**Introduction**

Ireland's forestry stock of 731,000 ha covers about 10.5% of the country, of which 44% is planted on peats (National Forest Inventory, 2012). Peats, especially ombrotrophic upland blanket peats, are generally lacking in minerals like aluminum (Al) and iron (Fe), and consequently have extremely low soil phosphorus (P) sorption capacities (Renou et al., 2000). As a result, any P released by the forestry operations, such as clearfelling and afforestation, is liable to leach unimpeded into adjacent receiving waters, even with the application of current best management practices targeted at preventing such pollution (Finnegan et al., 2014).

Nutrient enrichment, or eutrophication, of inland waters is recognized as Ireland’s most serious environmental pollution problem (Department of the Environment, 2002). As P has been identified as the primary nutrient limiting eutrophication in freshwaters (Carpenter et al., 1998; Boesch et al., 2001), preventing its migration from soil to aquatic environments is paramount. The oligotrophic nature of Ireland's upland catchments, and the unique flora and fauna present in these waters, make them particularly sensitive to eutrophication (Mainstone and Parr, 2002; Hutton et al., 2008). These forested peat catchments are headwaters for many of Ireland’s river systems, a great number of which contain important salmonid populations (Giller and O’Halloran, 2004), as well as other species protected under European Union (EU) legislation (O’Driscoll et al., 2012; Reid et al., 2013). Consequentially, pollution from diffuse, low concentration sources of P, such as forestry, is capable of causing considerable environmental damage to an area much larger than that which is forested.

The sustained release of P following forestry harvesting has been highlighted as an issue of particular concern, as much of Ireland's current peatland forestry stock was planted between the 1950s and the 1990s (Renou and Farrell, 2005) and has now reached, or is reaching,
harvestable age (Rodgers et al., 2010). Clearfelling is the harvesting technique most prevalent in Ireland, and accounted for 76.6% of timber felled between 2006 and 2012 (National Forest Inventory, 2012). Clearfelling involves the removal from site of the commercially-viable portions of the forestry crop (i.e. tree trunks), leaving large amounts of P to remain onsite, present both in the soil and in the non-commercial logging residues, or 'brash'. This brash accounts for a considerable percentage of the above ground nutrients contained in a typical coniferous tree (Moffat et al., 2006), and it has been shown to release these nutrients for many years following its deposition (Titus and Malcolm, 1999; Hyvönen et al., 2000). To compound the issue, the use of brash mats to form temporary driving surfaces for heavy felling machinery is an essential management practice, as it prevents serious damage to the underlying soil (Moffat et al., 2006). Clearfelling completely disrupts natural P cycling in a forest ecosystem, with the deposition of brash mats resulting in increased P availability, while the concurrent removal of trees from site results in decreased P uptake and sequestration - a situation which greatly increases the threat of P leaching to aquatic ecosystems (Schaller et al., 2015).

While rates of afforestation on blanket peats have been in decline in recent years (National Forest Inventory, 2012), the harvesting of presently established forest is inevitable, as are the resulting water quality issues, if effective pollution mitigation measures are not implemented. Current best management practices advocate the use of riparian buffer strips between forestry and adjacent aquatic zones (Forest Service, 2000). Previous research has found the use of suitably large riparian buffer zones on peat soils to be largely successful in protecting surface waters from influxes of suspended sediments and associated particulate nutrients (Nieminen et al., 2005). However, with peat's low P adsorption capacity, the performance of these buffer zones in satisfactorily mitigating the flow of dissolved P off site varies greatly. Moreover, it has been found that these buffer zones' effect can be anywhere from positive, with total
retention of released P (Vaaninen et al., 2008), to negative, compounding the issue with
additional P release (Vasander et al., 2003; Liljaniemi et al., 2003).

There has been increasing interest in the use of soil amendments to control P losses from
diffuse sources, such as land spreading of manure from dairy cattle (Brennan et al., 2011a,b)
and pigs (O’Flynn et al., 2012), land spreading of dairy waste water from washing of milking
parlors (Serrenho et al., 2012), and construction of wetlands on sites previously used for
agriculture (Ann et al., 1999). The use of chemical amendments has shown much promise in
these instances, though there has been little to no investigation into the practice’s potential in
abating the loss of P from peatland forestry. The performance of chemical amendments in
this context has the potential to be quite different, as the interaction of amendments with the
complex chemistry of peat can significantly alter their P removal performance (James et al.,

This study tested the hypothesis that chemical amendments could immobilize P in a forest
peat soil. To address this hypothesis, the aim of this study was to identify and compare
potential soil amendments which function well in a peat chemistry environment and, when
mixed with peat onsite, could increase the adsorptive capacity of the soil to the point where P
losses from a forestry site would no longer pose a risk to receiving surface waters.
Specifically, the objectives of this study were to: (1) identify chemical amendments capable
of removing P from an aqueous solution which mimics the chemistry of runoff/ground water
on a forested peat site (2) compare the performance of each amendment and identify the
amendment most ideally suited to application in a forested peat site (3) analyze the effect of
peat on the performance of each amendment in terms of its adsorptive capacity.
Materials and Methods

Collection and characterization of peat samples

Samples of blanket peat were collected from a recently clearfelled riparian buffer zone, located near the town of Oughterard, County Galway, Ireland (approx. coordinates 53.375N, -9.419E). The peat was collected from a stratum of homogeneous composition, at a depth of 5-25cm from the surface (the surface layer of vegetation and semi-decayed sphagnum moss was discarded, along with any larger roots and plant fibers). Before testing, the samples were homogenized by repeatedly folding and kneading the peat until its texture and water content were entirely uniform. Once homogenized, peat samples were sealed in airtight Ziploc® bags and stored in a temperature-controlled room at 11°C until testing commenced. The water content of the homogenized peat was determined to be 89.7%±0.08% by oven drying the peat for 24 h at 105°C (BSI, 1990).

Sourcing and characterization of amendments

The following amendments were used in this study (Table 1): dewatered aluminum sulphate water treatment residual (Al-WTR; oven dried for 24 h at 105°C and ground to pass a 0.5 mm sieve), crushed concrete cubes (pulverized with a mortar and pestle and ground to pass a 0.5 mm sieve), gypsum (sourced from recycled plasterboard, supplied as a powder, and passed through a 0.5 mm sieve to ensure uniformity), magnesium oxide (MgO; sourced from seawater, supplied as a powder, and passed through a 0.5 mm sieve to ensure uniformity), magnesium hydroxide (Mg(OH)₂; sourced from seawater, supplied as a paste, and passed through a 0.5 mm sieve to ensure uniformity) and steel wool (grade 00, fiber diameter approx. 0.04 mm).

Batch test procedure
Samples were prepared comprising two grams (wet weight) of the homogenized peat, mixed with each of the six amendments, at amendment to peat-solid ratios of 1:10, 1:4 and 1:2. These masses provided a material to solution ratio that was small enough to ensure non-zero equilibrium concentrations, thus allowing the determination of the entire isotherm curve. The material to solution ratio used in the current study was comparable to ratios used in similar studies (Li et al., 2006; Chardon et al., 2012). The samples were placed in separate 50 ml conical flasks and overlain by 25 ml of deionized water, with ortho-phosphorus (PO$_4$-P) concentrations (prepared by adding various amounts of K$_2$HPO$_4$ to deionized water) of 0, 25, 50, 75, 100, and 150 mg P L$^{-1}$. This range of P concentrations was determined to be sufficiently wide to account for the variation in the amendments’ adsorptive capacities, and the dependence of adsorption capacity on the initial concentration (Seo et al., 2005). The flasks were sealed with Parafilm and shaken in a reciprocal shaker (250 rpm) at room temperature (25°C) for 24 h. After 24 h had elapsed, the samples were centrifuged at 14,500 rpm for 5 min, and the supernatant water was passed through a 0.45 µm filter. Dissolved P concentrations of the supernatant water were determined using a nutrient analyzer (Konelab 20, Thermo Clinical Lab systems, Finland) after APHA (1998). Experiments were conducted at neutral pH to ensure that all observed changes in pH could be attributed to the addition of peat to the solution. The same procedure was also carried out using identical masses of amendment but without the addition of peat.

**Analysis of experimental data**

The mass of P adsorbed per gram of adsorbent, $q_e$, at equilibrium was calculated by:

$$q_e = \frac{V(C_0 - C_e)}{m}$$  \hspace{1cm} (1)

where $C_0$ and $C_e$ are the initial and final PO$_4$-P concentrations of the solution (mg L$^{-1}$), $V$ is the volume of solution (L), and $m$ is the dry weight of adsorbent (g). For the purpose of these
calculations, the PO₄-P adsorbency of the peat was considered to be negligible, and for
solutions containing peat and amendment, only the dry weight of the amendment was
considered in Eqn 1.

The Ce and qₑ data obtained from the batch tests were fitted to the Koble-Corrigan equation,
which was chosen for its ability to model adsorption over a wide range of data, as well as for
its being more universally applicable than the commonly used Langmuir and Freundlich
equations (Koble and Corrigan, 1952). The Koble-Corrigan equation is as follows:

\[ qₑ = \frac{A_{KC}Cₑ^p}{1 + B_{KC}Cₑ^p} \]  (2)

where \( A_{KC}, B_{KC} \) and \( p \) are the Koble-Corrigan isotherm constants determined by using an
iterative approach to minimize the value returned by the hybrid fractional error function
(HYBRID), an error function used to fit Eqn. 2 to the observed data (Porter et al., 1999):

\[ \frac{100}{n-p} \sum_{i=1}^{n} \frac{(qₑ,i,meas-qₑ,i,calc)^2}{qₑ,i,meas} \]  (3)

where \( n \) is the number of experimental data points, \( p \) is the number of isotherm constants in
the Koble-Corrigan equation (Eqn. 2), \( qₑ,i,meas \) is the value of \( qₑ \) obtained from Eqn. (1), and
\( qₑ,i,calc \) is the value of \( qₑ \) obtained from Eqn. (2). By fitting the experimental data for each
adsorbent to the Koble-Corrigan equation, it was possible to make predictions for values of
\( Cₑ \) and \( qₑ \) for any given \( Cₒ \). Substituting Eqn. 1 into Eqn. 2 gives:

\[ \frac{V(cₒ-Cₑ)}{m} = \frac{A_{KC}Cₑ^p}{1 + B_{KC}Cₑ^p} \]  (4)

An iterative approach, using Microsoft Excel's Solver™, was then used to determine values
of \( Cₑ \) for any given value of \( Cₒ \). The impact of peat on the amendments’ performance was
evaluated using:
where $q_{e,a}^{\text{calc}}$ is the modelled mass of P adsorbed by the amendment in a solution containing peat, and $q_{e,b}^{\text{calc}}$ is the modelled mass of P adsorbed by the amendment in a solution containing no peat. Calculated values of $\Phi > 1$ indicate a synergistic effect was obtained by exposing an amendment to peat, i.e. the influence of peat chemistry was favorable for P adsorption, while values of $\Phi < 1$ suggest the opposite, i.e. that the effect was antagonistic. In this way, $\phi$ could be considered to be a coefficient of synergy.
Results and Discussion

All of the amendments were effective at removing P from aqueous solution, both with and without the presence of peat in solution. The addition of peat did, however, influence the behavior of all amendments, leading to either improved or diminished adsorptive performance.

Equilibrium adsorption isotherms

Figure 1 shows the adsorption isotherms obtained by fitting the Koble-Corrigan model to each of the peat amendment-mixtures. The Koble-Corrigan fitting coefficients (Akc, Bkc, and p) as well as slope, and R² values obtained from plots of qe calc (calculated using the Koble-Corrigan equation) vs qe exp, are shown in Table 2. The Koble-Corrigan model fitted the experimental data well, and there was a very good correlation (average R² = 0.94±0.1) between predicted and experimental values. In general, higher values of qe were observed when the ratio of adsorbent to solution was smaller. This is in agreement with a review conducted by Cucarella and Renman (2009), who reported that smaller adsorbent-to-solution ratios may lead to higher concentrations of P sorbed to the adsorbent material. Al-WTR, in particular, showed a much larger increase in P adsorption capacity (maximum observed qe = 11.8 mg g⁻¹) when added to peat-containing solutions at an amendment to peat-solid ratio of 1:10, compared to amendment to peat-solid ratios of 1:4 and 1:2 (maximum observed qe = 3.3 mg g⁻¹ and 3.6 mg g⁻¹, respectively).

Maximum qe values observed for concrete were 15 mg g⁻¹, 8.9 mg g⁻¹, and 6.8 mg g⁻¹ at amendment to peat-solid ratios of 1:10, 1:4, and 1:2, which were similar to qe values of 26.8 mg g⁻¹, 15.1 mg g⁻¹, and 6.6 mg g⁻¹ observed for the same masses of concrete, but without the addition of peat. These values are all comparable to the qe range of 5.1-19.6 mg g⁻¹ observed by Egemose et al. (2012). With respect to gypsum, the presence of peat in solution resulted in
greatly reduced maximum observed $q_e$ values of 10.8 mg g$^{-1}$, 4.4 mg g$^{-1}$, and 3.6 mg g$^{-1}$ at amendment to peat-solid ratios of 1:10, 1:4, and 1:2, compared to values of 84.8 mg g$^{-1}$, 58.4 mg g$^{-1}$, and 34 mg g$^{-1}$ observed for the same masses without the addition of peat.

Magnesium oxide and Mg(OH)$_2$ displayed the highest P adsorption capacities, with MgO having the greatest of the two, with a maximum $q_e$ of 102.7 mg g$^{-1}$ observed at an amendment to peat-solid ratio of 1:10. Liu et al. (2011) found that removal of As(III) from aqueous solution by MgO was due to the *in situ* formation of Mg(OH)$_2$ by reaction of the MgO with water, followed by subsequent adsorption/reaction of the newly formed Mg(OH)$_2$ with the As(III) anion. Liu et al. (2011) also reported that the adsorptive performance of Mg(OH)$_2$ formed *in situ* was greater than that of pre-formed Mg(OH)$_2$, at least partially as a result of the former’s larger specific surface area, which was almost 5.5 times greater than that of the latter (58.4 m$^2$g$^{-1}$ vs. 10.7 m$^2$g$^{-1}$, respectively). The results of our study indicate that Mg(OH)$_2$ formed *in situ* also had a greater adsorptive capacity for PO$_4$–P compared to that of pre-formed Mg(OH)$_2$. Stoichiometrically, MgO contains approximately 45% more Mg per unit weight than Mg(OH)$_2$; however, its P adsorption capacity was observed to be up to 120% greater than that of preformed Mg(OH)$_2$ in solutions where peat was absent (data not shown).

When steel wool was examined, P adsorption was observed to be inhibited at high initial PO$_4$–P concentrations ($C_0$), both with and without the addition of peat. This appeared to be the result of high PO$_4$–P concentrations inhibiting the formation of the iron oxides/hydroxides responsible for P removal, and it was observed that much less of the steel wool had rusted in these comparatively high concentration solutions. Similar observations were made by Pryor and Cohen (1953), who found that solutions of orthophosphate passivate iron; i.e. iron became less reactive due to presence of a micro coating in the presence of dissolved air, and
Harahuc et al. (2000), who found that phosphate at or above concentrations of 25 mmol inhibited the solubilization of iron.

**pH effects**

Figure 2 summarizes the effect the addition of peat had on the pH of the solution at equilibrium (after 24 h). The peat displayed a strong buffering capacity, and its addition altered the final pH of the solution at equilibrium in all cases. In real-world applications, an amendment that has a strong effect on pH may not be desirable, as large fluctuations in pH may have a deleterious impact on the environment. For example, the addition of an amendment which increases the soil’s pH may increase soil nitrate (NO₃) concentrations, and the risk of NO₃ leaching could be particularly great during the clearfelling of forestry (Wickström, 2002). After shaking for 24h, the pH of samples containing peat mixed with Al-WTR was found to be closest to those of a solution containing peat only. The performances of P adsorbent materials are known to be strongly affected by pH, and in acidic soils such as peats, Al and Fe are known to be responsible for the sorption of P (Sato et al., 2005), while binding by calcium (Ca) and magnesium (Mg) is responsible for P immobilization in alkaline soils (Faulkner and Richardson, 1989; Reddy et al., 1999).

**Synergistic Effects**

The results indicate that, at certain initial PO₄–P concentrations, a degree of synergy exists between the peat and certain amendments; i.e. their adsorption performance was better in an environment whose chemistry was influenced by peat (Figure 3). Al-WTR and steel wool, in particular, showed improved performance across a large range of concentrations, while concrete, gypsum, and MgO showed largely diminished efficacy. The effect of peat on P adsorption by Mg(OH)₂ varied from beneficial to detrimental, depending on the ratio of peat to amendment in solution.
One mechanism by which an amendment's P sorption performance may improve when in a peat environment is as a result of the interaction of humic substances with soluble metal ions. Humic substances facilitate the removal of P through reactions with Al and Fe ions, which results in the formation of insoluble phosphate complexes (Bloom, 1981; Renou et al., 2000). Conversely, humic materials inhibit the removal of P by precipitation with Ca by competing with the phosphate anion for Ca, with this effect being more exaggerated at lower pH (Alvarez et al., 2004; Song et al., 2006; Cao et al., 2007). The formation of humic complexes may therefore contribute both to the observed improvement in Al-WTR and steel wool’s adsorptive performance, as well as the diminished performance of concrete and gypsum. Phosphorus sorption by gypsum, for example, is directly related to its solubility, as its P removal is largely due to precipitation reactions (Penn et al., 2007) and thus the availability of Ca for the formation of calcium phosphates (Kordlaghari and Rowell, 2006).

The concrete used in this study had significantly higher levels of Al and Fe present than the gypsum (Table 1), which may explain why its P removal performance was not as adversely affected as gypsum's. Öğuz et al. (2003) found that P removal by concrete was by precipitation of metallic PO₄ salts, which were adsorbed onto the surface of the concrete, and AlPO₄ appeared to be the main product adsorbed onto the surface of the gas concrete examined in their study. This result suggests that the Al and Fe content of the concrete examined in this study, perhaps more so than its Ca content, plays a crucial role in P adsorption. Berg et al. (2005) reported that P removal by crushed gas concrete was not affected by the presence of organic matter in the form of dissolved organic carbon. However, as concrete’s adsorption of P is strongly related to pH, and improves with increasing pH (Öğuz et al., 2003), it appears that the acidification of the solution by the peat is at least partially responsible for the observed reduction in performance.
The pH buffering action of the peat is also likely to have contributed to the observed synergistic effects by lowering the solution pH, which can improve Al-WTR’s P adsorption, as observed by Yang et al. (2006), and promote corrosion of the steel wool to form P adsorbing iron oxide/hydroxides. Consequently, the peat had a stronger effect on pH at this rate, with the final solution pH ranging from 8.3-8.6, compared to 9.8-10.2 and 9.9-10.5 at amendment to peat-solid ratios of 1:4 and 1:2, respectively. It was also observed that when Mg(OH)₂ was added at an amendment to peat-solid ratio of 1:10, recorded qₑ values were much lower than solutions containing Mg(OH)₂ at amendment to peat-solid ratios of 1:4 and 1:2. While the reason for this is not entirely clear, the marked difference in pH accompanied by the significant drop in P adsorption strongly suggests that the pH buffering effect of the peat may also be responsible for reducing the P adsorbency of Mg(OH)₂ at this lowest rate of amendment. Xie et al. (2013) reported that fulvic and humic acids had a slight negative effect on PO₄ adsorption by a magnesium oxide nanoflake-modified diatomite adsorbent, and attributed this observation to the effects of competitive adsorption. It is quite possible that at low Mg(OH)₂ to peat ratios this effect is more exaggerated.
Conclusions

This study found that the presence of peat in solution increased the P adsorption capacities of Al-WTR and steel wool, while the adsorptive capacities of crushed concrete, gypsum, and magnesium oxide were decreased. Magnesium hydroxide showed both increased and decreased adsorptive capacity, depending on the ratio of peat to amendment in solution. Throughout the study, peat displayed considerable ability to buffer the peat-amendment mixture pH. However, the equilibrium pH of solutions containing Al-WTR and gypsum were observed to be closest to the peat's natural pH, indicating that these amendments were best suited to the peat environment. Taking these factors into consideration, the results of this study indicate that Al-WTR, in particular, holds great promise for utilization in the mitigation of P runoff caused by forestry operations on blanket peat sites. Previous research into the use of Al-WTR has shown it to be effective in preventing P runoff when used as a soil amendment in an agricultural context, and this study's findings demonstrate that the chemistry of a peat forest environment is likely to interact with Al-WTR in such a manner as to improve its P removal performance at low P concentrations. As a waste material, an additional benefit lies in the fact that it does not have the production costs associated with other high performing amendments studied, i.e. steel wool and Mg(OH)₂.

Acknowledgements

The first author would like to acknowledge the Irish Research Council for funding.
References


Legend

Figure 1. Best fit Koble-Corrigan adsorption isotherm models derived using non-linear regression methods for (a) magnesium oxide, (b) gypsum, (c) steel wool, (d) crushed concrete, (e) magnesium hydroxide, and (f) Al-WTR, at amendment to peat-solid ratios of 1:10, 1:4, and 1:2.

Figure 2. Effect of peat on pH at equilibrium

Figure 3. Synergistic effect of peat on P adsorption (Refer to Eqn. 5)
Table 1. Peat and amendment characteristics

<table>
<thead>
<tr>
<th>Material</th>
<th>Total exchange capacity (meq 100 g⁻¹)</th>
<th>pH</th>
<th>Organic matter (%)</th>
<th>P† (mg kg⁻¹)</th>
<th>Bray II P (mg kg⁻¹)</th>
<th>Ca† (mg kg⁻¹)</th>
<th>Mg† (mg kg⁻¹)</th>
<th>Fe† (mg kg⁻¹)</th>
<th>Al† (mg kg⁻¹)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Peat</td>
<td>11.45</td>
<td>3.8</td>
<td>100</td>
<td>5</td>
<td>3</td>
<td>228</td>
<td>142</td>
<td>152</td>
<td>47</td>
</tr>
<tr>
<td>Al-WTR</td>
<td>16.78</td>
<td>7</td>
<td>21.62</td>
<td>2</td>
<td>&lt; 1</td>
<td>2983</td>
<td>32</td>
<td>154</td>
<td>2159</td>
</tr>
<tr>
<td>Gypsum</td>
<td>64.3</td>
<td>8.8</td>
<td>1.85</td>
<td>6</td>
<td>14</td>
<td>11858</td>
<td>296</td>
<td>111</td>
<td>66</td>
</tr>
<tr>
<td>Crushed Concrete</td>
<td>145.15</td>
<td>11.7</td>
<td>0.82</td>
<td>&lt; 1</td>
<td>&lt; 1</td>
<td>27593</td>
<td>622</td>
<td>405</td>
<td>441</td>
</tr>
<tr>
<td>MgO</td>
<td>267.55</td>
<td>10</td>
<td>1.04</td>
<td>&lt; 1</td>
<td>&lt; 1</td>
<td>2734</td>
<td>29760</td>
<td>77</td>
<td>&lt; 1</td>
</tr>
<tr>
<td>Mg(OH)₂</td>
<td>247.78</td>
<td>9.8</td>
<td>23.53</td>
<td>&lt; 1</td>
<td>&lt; 1</td>
<td>739</td>
<td>28660</td>
<td>57</td>
<td>&lt; 1</td>
</tr>
</tbody>
</table>

†Mehlich-3 extractable
Table 2. Koble-Corrigan fitting coefficients, HYBRID errors, and values of slope and $R^2$ obtained from plots of $q_e$ calc vs $q_e$ exp for all amendments studied.

<table>
<thead>
<tr>
<th>Amendment</th>
<th>Mass of amendment (g)</th>
<th>Mass of Peat (g)</th>
<th>Koble-Corrigan Parameters</th>
<th>Hybrid Error</th>
<th>$q_e$ calc vs $q_e$ exp</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td></td>
<td>$A_{kc}$</td>
<td>$B_{kc}$</td>
<td>$P$</td>
</tr>
<tr>
<td>Al-WTR</td>
<td>0.021</td>
<td>-</td>
<td>0.004</td>
<td>0.000</td>
<td>1.748</td>
</tr>
<tr>
<td></td>
<td>0.051</td>
<td>-</td>
<td>0.002</td>
<td>0.001</td>
<td>1.844</td>
</tr>
<tr>
<td></td>
<td>0.103</td>
<td>-</td>
<td>0.029</td>
<td>0.010</td>
<td>1.090</td>
</tr>
<tr>
<td></td>
<td>0.021</td>
<td>2.000</td>
<td>0.710</td>
<td>0.032</td>
<td>0.795</td>
</tr>
<tr>
<td></td>
<td>0.051</td>
<td>2.000</td>
<td>0.669</td>
<td>0.259</td>
<td>1.201</td>
</tr>
<tr>
<td></td>
<td>0.103</td>
<td>2.000</td>
<td>0.001</td>
<td>-0.999</td>
<td>0.000</td>
</tr>
<tr>
<td>Crushed Concrete</td>
<td>0.021</td>
<td>-</td>
<td>1.031</td>
<td>-0.072</td>
<td>0.446</td>
</tr>
<tr>
<td></td>
<td>0.051</td>
<td>-</td>
<td>0.510</td>
<td>0.020</td>
<td>0.873</td>
</tr>
<tr>
<td></td>
<td>0.103</td>
<td>-</td>
<td>1.929</td>
<td>0.205</td>
<td>0.514</td>
</tr>
<tr>
<td></td>
<td>0.021</td>
<td>2.000</td>
<td>3.964</td>
<td>-0.099</td>
<td>0.210</td>
</tr>
<tr>
<td></td>
<td>0.051</td>
<td>2.000</td>
<td>1.087</td>
<td>0.109</td>
<td>0.935</td>
</tr>
<tr>
<td></td>
<td>0.103</td>
<td>2.000</td>
<td>1.625</td>
<td>0.169</td>
<td>0.585</td>
</tr>
<tr>
<td>Gypsum</td>
<td>0.021</td>
<td>-</td>
<td>0.191</td>
<td>0.002</td>
<td>1.722</td>
</tr>
<tr>
<td></td>
<td>0.051</td>
<td>-</td>
<td>0.004</td>
<td>0.000</td>
<td>2.841</td>
</tr>
<tr>
<td></td>
<td>0.103</td>
<td>-</td>
<td>0.089</td>
<td>-0.950</td>
<td>0.016</td>
</tr>
<tr>
<td></td>
<td>0.021</td>
<td>2.000</td>
<td>0.504</td>
<td>-0.381</td>
<td>0.176</td>
</tr>
<tr>
<td></td>
<td>0.051</td>
<td>2.000</td>
<td>0.000</td>
<td>-0.999</td>
<td>0.000</td>
</tr>
<tr>
<td></td>
<td>0.103</td>
<td>2.000</td>
<td>0.000</td>
<td>-0.999</td>
<td>0.000</td>
</tr>
<tr>
<td>Magnesium Oxide</td>
<td>0.021</td>
<td>-</td>
<td>1.686</td>
<td>0.012</td>
<td>2.406</td>
</tr>
<tr>
<td></td>
<td>0.051</td>
<td>-</td>
<td>27.452</td>
<td>0.442</td>
<td>1.438</td>
</tr>
<tr>
<td></td>
<td>0.103</td>
<td>-</td>
<td>17.994</td>
<td>0.252</td>
<td>0.520</td>
</tr>
<tr>
<td></td>
<td>0.021</td>
<td>2.000</td>
<td>23.555</td>
<td>0.177</td>
<td>0.827</td>
</tr>
<tr>
<td></td>
<td>0.051</td>
<td>2.000</td>
<td>24.286</td>
<td>0.407</td>
<td>1.197</td>
</tr>
<tr>
<td></td>
<td>0.103</td>
<td>2.000</td>
<td>18.636</td>
<td>0.895</td>
<td>0.774</td>
</tr>
<tr>
<td>Magnesium Hydroxide</td>
<td>0.021</td>
<td>-</td>
<td>1.214</td>
<td>-0.159</td>
<td>0.379</td>
</tr>
<tr>
<td></td>
<td>0.051</td>
<td>-</td>
<td>0.055</td>
<td>0.001</td>
<td>1.474</td>
</tr>
<tr>
<td></td>
<td>0.103</td>
<td>-</td>
<td>0.152</td>
<td>-0.939</td>
<td>0.011</td>
</tr>
<tr>
<td></td>
<td>0.021</td>
<td>2.000</td>
<td>0.047</td>
<td>-0.009</td>
<td>1.011</td>
</tr>
<tr>
<td></td>
<td>0.051</td>
<td>2.000</td>
<td>7.849</td>
<td>0.098</td>
<td>1.065</td>
</tr>
<tr>
<td></td>
<td>0.103</td>
<td>2.000</td>
<td>2.075</td>
<td>0.047</td>
<td>1.473</td>
</tr>
<tr>
<td>Steel wool</td>
<td>0.021</td>
<td>-</td>
<td>2.921</td>
<td>0.159</td>
<td>1.484</td>
</tr>
<tr>
<td></td>
<td>0.051</td>
<td>-</td>
<td>4.128</td>
<td>0.411</td>
<td>0.996</td>
</tr>
<tr>
<td></td>
<td>0.103</td>
<td>-</td>
<td>1.181</td>
<td>0.156</td>
<td>1.001</td>
</tr>
<tr>
<td></td>
<td>0.021</td>
<td>2.000</td>
<td>434.832</td>
<td>9.756</td>
<td>1.352</td>
</tr>
<tr>
<td></td>
<td>0.051</td>
<td>2.000</td>
<td>41.715</td>
<td>0.926</td>
<td>1.313</td>
</tr>
<tr>
<td></td>
<td>0.103</td>
<td>2.000</td>
<td>68.370</td>
<td>2.490</td>
<td>0.728</td>
</tr>
</tbody>
</table>
Figure 1.

Magnesium Oxide

Gypsum

Steel Wool

Crushed Concrete

Magnesium Hydroxide

Alum Water Treatment Residual

Ratio of amendment to peat-solids

1:10  
1:4  
1:2
Figure 2.

Mass of amendment
- 0.021g
- 0.051g
- 0.103g

No peat in solution
- ■
- ○
- ▲

Peat in solution
- ■
- ○
- ▲

Peat without amendment
- —
Figure 3.

<table>
<thead>
<tr>
<th>Material</th>
<th>Initial P Concentration (mg L⁻¹)</th>
<th>Ratio of amendment to peat-solids</th>
</tr>
</thead>
<tbody>
<tr>
<td>Magnesium Oxide</td>
<td>25, 50, 75, 100, 125, 150</td>
<td>1:10</td>
</tr>
<tr>
<td>Gypsum</td>
<td>25, 50, 75, 100, 125, 150</td>
<td>1:4</td>
</tr>
<tr>
<td>Steel Wool</td>
<td>25, 50, 75, 100, 125, 150</td>
<td>1:2</td>
</tr>
<tr>
<td>Concrete</td>
<td>25, 50, 75, 100, 125, 150</td>
<td></td>
</tr>
<tr>
<td>Magnesium Hydroxide</td>
<td>25, 50, 75, 100, 125, 150</td>
<td></td>
</tr>
<tr>
<td>Alum Water Treatment Residual</td>
<td>25, 50, 75, 100, 125, 150</td>
<td></td>
</tr>
</tbody>
</table>