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A F-bridged Mn(II) Molecular Square

Seán T. Meally, Kevin Mason, Patrick McArdle, Euan. K. Brechin, Alan G. Ryder, and Leigh F. Jones*

We report on a tetranuclear Mn(II) complex possessing a F-bridged molecular square topology directed by 1,10-phenanthroline ligands. A significant rise in the use of F-bridging ions in the field of molecular magnetism has been observed over the last 5-10 years. Perhaps the most elegant example is in the synthesis of a vast library of homo- and heterometallic \{CrIII\n\textit{7}MII\textit{1}\} wheels (and their many analogues) reported by Winpenny and co-workers, that have been proposed as qubits in quantum information processing (QIP). The bridging F- ions encircling the inner cavity of the wheels H-bond to templating alkylammonium cations, changes in which allow for modification of both the molecular and supramolecular chemistry; a beautiful example being the formation of inorganic-organic rotaxanes. Such dipolar H-F interactions have also been used to manipulate the docking of ‘guest’ molecules (inc. Na+, K+, NH4+, β-Alanine, MeNH3+) within the double bowl cavities of hexanuclear F-bridged [Cu6] wheels. F-bridged cages of MnII/III are far rarer in comparison to other 3d metals, perhaps because of the perceived lack of suitable starting materials. MnF2 was used for the first time very recently for the preparation of a host of polymetallic high spin complexes and Single-Molecule Magnets (SMMs); its prior lack of use a surprise given the vast majority of SMMs are built using MnIII ions. In order to further the coordination chemistry of F- with Mn, and with an eye towards investigating Mn-F-Mn magnetic exchange interactions and developing magneto-structural correlations within various low nuclearity F-bridged complexes, we herein describe an unusual molecular square built from MnF2. Reaction of anhydrous MnF2, 1,10-phenanthroline (phen) and NaNO3 in MeOH† for 24 hours produces a yellow solution, which upon filtration and slow evaporation, yields yellow block-like crystals ‡ of \[\text{[Mn}\text{II}\text{4F4(phen)8]}(\text{NO3})_4\cdot12\text{H2O} \ (1)\], crystallising in the tetragonal I4 1/a space group in ~20% yield (Figure 1). The structure comprises four symmetry equivalent \{Mn(phen)2\}2+ units (Mn1) linked into a square by four µ-bridging F- ions (F1 and s.e.). The Mn ions are in distorted octahedral geometries with a \{N 4\}F2 coordination sphere. The two chelating phen ligands and two fluorides per metal centre are both cis, directing the molecular square topology (Fig. 1) - analogous to that observed, for example, in PdII chemistry. As can be seen in the lower panel of Figure 1 the \[\text{[Mn}\text{II}\text{4F4]}^4+\] core is non-planar, being bent along the body diagonal. Each edge of the square is 3.722 Å (Mn…Mn) in length, with the Mn-F bond distances being Mn1-F1 = 2.061(4) Å and Mn1-F′ = 2.030(4) Å. The µ-F bridges sit alternately above and below the \{Mn4\} “plane” as you circumnavigate the square. The combination of the non-linear Mn-F-Mn pathways (Mn1-F1-Mn1, 130.9(2)º) and the presence of the bulky phen ligands are presumably the cause of the molecular distortion. Intra-molecular interactions in the form of π-π stacking of the phen ligands are observed in the lattice with πcentroid…πcentroid distances in the 3.719 – 3.886 Å range (Fig. 1). The individual [Mn4F4(phen)8]+ moieties align in superimposable 1D rows along the c axis of the unit cell although they are well separated ([Mn4] plane-plane distance = 13.231 Å). The NO3- counter anions (O1, O2 and O3) and the H2O of crystallisation (O15, O16 and O17) lie in between these 1D columns and H-bonding with each other (O1-O16 = 2.839, O2-O15 = 2.779, O2-O17 = 2.890 Å), and with phen protons of nearby [Mn4] units (O1-H3(C3) =2.588, O2-H8(C8) = 2.590, O15-H9(C9) = 2.563 Å). The 1D columns (Fig. 2) interact via numerous π-π stacking interactions (πcentroid…πcentroid distance = 3.654 Å, Fig. 2).
Fig. 1 Molecular structure of the cation of 1 viewed perpendicular (top) and parallel (bottom) to the [Mn4] plane. H atoms have been omitted for clarity. Mn = pink; F = green; N = blue, C = grey.

Oxidation state assignments and the presence of F (versus, for example, OH) were made on the basis of bond lengths, BVS calculations (detailed in ESI), charge balance considerations and Raman spectroscopy (vide infra). Although polymetallic complexes with square topologies are not uncommon, the synthesis of 1 represents the first reported F-bridged molecular square.

Fig. 2 (Top) Packing of cations of 1 in the crystal (viewed down c axis) highlighting the 1D columnar arrangement of [Mn4] units. Arrows represent examples of intermolecular π Stacking in 1. (bottom) View along the b axis of the unit cell. In both cases the H-atoms, the NO3- counter anions and the H2O solvent molecules have been removed for clarity.

Variable temperature magnetic susceptibility measurements were carried out on a polycrystalline powder of 1 in the 300 – 5 K temperature range in an applied field of 0.1 T (Fig. 3). The room temperature χmT value of ~12.8 cm3 mol-1 K is significantly lower than that expected for four non-interacting MnII centres (~17.5 cm3 mol-1 K for g = 2.0) and is indicative of antiferromagnetic exchange between the four metal centres. The value of χmT drops gradually with decreasing temperature before falling more rapidly below ~125 K reaching a value of 0.47 cm3 mol-1 K at 5 K. Simulation of the magnetic data10 using the 1-J model of Equation (1) and the scheme of Figure 3 afforded the parameters J = -3.3 cm-1 and g = 2.00; the magnitude of the exchange being in-line with previously reported values for MnII-F-MnII interactions.11

\[ H = -2J(S_1 \cdot S_2 + S_3 \cdot S_4 + S_2 \cdot S_3 + S_1 \cdot S_4) \] (1)

Solid state Raman spectra were collected on a polycrystalline sample of 1 and the MnF2 starting material (purchased as a powder) in order to confirm the presence of the μ2-F- bridging ions (Figure 4). Peaks at 242 and 254 cm-1 attributed to the Mn(II)-F-Mn(II) symmetric stretching modes in 1 were observed.12 The Raman spectrum of MnF2 exhibits symmetric stretch peaks at 242 and 245 cm-1 (Eg mode)13 and a peak at 336 cm-1 consistent with terminal Mn-F stretching (A1g mode)13 which is corroborated by its absence in the spectrum of 1 (Fig. 3). The band at 273 cm-2 is due to the organic ligand.

The electrospray mass spectrum gives a prominent parent peak at m/z = 434 ([Mn4F4(phen)8]4+), which suggests that 1 remains intact in solution, although this may also be attributed to the [MnF(phen)2]4+ fragment species. Other smaller fragments of 1 were also observed as assigned in Fig. SI3. UV/vis studies on 1 in MeCN show two sharp bands (226 and 266 nm respectively),
and stirred for 24 h before being filtered to give a yellow solution. Yellow ions no d-d transitions were observed (Fig. SI4). In conclusion, as expected for isotropic Mn(II) transitions was also observed. As expected for isotropic Mn(II) with an exposure time of 3 × 5 seconds was used and spectra were recorded at room temperature using a Raman spectrometer out at the School of Chemistry, NUI Galway. Raman measurements were collected from 250 to 3311 cm⁻¹ (at a resolution of 1 cm⁻¹).

**Notes and References**


We report on the solid state and solution characterisation of a tetranuclear Mn(II) cluster possessing a unique F-based ‘molecular square’ core.